

CHAPTER 22 STUDY GUIDE FOR CONTENT MASTERY

Hydrocarbons

Section 22.1 Alkanes

In your textbook, read about organic chemistry, hydrocarbons, and straight-chain alkanes.

Use each of the terms below just once to complete the passage.

hydrocarbons

homologous series

organic compounds

straight-chain alkanes

Most compounds that contain carbon are known as **(1)** _____. The simplest group of such compounds are **(2)** _____, which contain only carbon and hydrogen. If all of the carbon atoms are linked by single covalent bonds and there are no branches, the compounds are called **(3)** _____. Ethane, propane, and butane are three examples. They are members of one **(4)** _____ because they differ from each other by a repeating unit ($-\text{CH}_2-$).

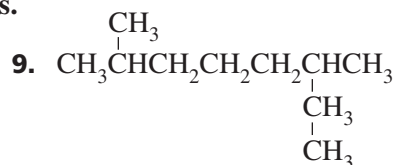
In your textbook, read about branched-chain alkanes and naming them.

For each statement below, write *true* or *false*.

- _____ 5. The ability of carbon atoms to bond to two, three, or four other carbon atoms makes possible a variety of branched-chain alkanes.
- _____ 6. A carbon atom or group of carbon atoms that branch off the main hydrocarbon chain of an alkane is a substituent group.
- _____ 7. A skeletal formula is a way of representing an organic compound by showing only the hydrogen atoms.

Use the IUPAC rules to name the following structures.





Draw the structure of each of the following alkanes.

10. 2-methylheptane

11. 2,3,4-trimethylpentane

Section 22.2 Cyclic Alkanes and Alkane Properties

In your textbook, read about cycloalkanes.

For each item in Column A, write the letter of the matching item in Column B.

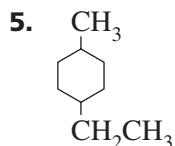
Column A

- _____ 1. A simplified way of representing an organic compound by showing only the carbon-carbon bonds
- _____ 2. A way of representing an organic compound that saves space by not showing how the hydrogen atoms branch off the carbon atoms
- _____ 3. Indicates that a hydrocarbon has a ring structure
- _____ 4. A hydrocarbon that has a ring of carbon atoms in its structure

Column B

- a. *cyclo-*
- b. condensed structural formula
- c. line structure
- d. cyclic hydrocarbon

Use the IUPAC rules to name the following structure.



Draw the structure of the following cycloalkane.

6. 1-ethyl, 2-propylcyclobutane

In your textbook, read about the properties of alkanes and multiple carbon-carbon bonds.

In the space at the left, write the word or phrase in parentheses that correctly completes the statement.

- _____ 7. All the bonds in an alkane are (polar, nonpolar).
- _____ 8. The attractive forces between alkane molecules are (stronger, weaker) than the attractive forces between alkane and water molecules.
- _____ 9. Alkanes are (very, not very) soluble in water.
- _____ 10. The boiling points of alkanes (increase, decrease) with increasing molecular mass.
- _____ 11. The chief chemical property of alkanes is their (low, high) reactivity.
- _____ 12. Alkanes are often used as (solvents, fuels) because they readily undergo combustion in oxygen.
- _____ 13. Alkanes are (saturated, unsaturated) hydrocarbons because they have only single bonds.

Section 22.3 Alkenes and Alkynes

In your textbook, read about alkenes, alkynes, and naming alkynes.

Use the following words to complete the statements.

alkene	alkyne	electron density	ethene	ethyne
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- An _____ is a hydrocarbon that has one or more triple covalent bonds between carbon atoms.
- The unsaturated hydrocarbon _____ is the starting material for the synthesis of the plastic polyethylene.
- An _____ is a hydrocarbon that has one or more double covalent bonds between carbon atoms.
- Torches used in welding burn _____, which is commonly called acetylene.
- Alkenes and alkynes are more reactive than alkanes because double and triple bonds have greater _____ than single bonds have.

Circle the letter of the correct name for each of the following structures.

- | | |
|---|--|
| <p>6. $\text{CH}_3\text{CH}_2\text{CH}=\text{CHCH}_2\text{CH}_2\text{CH}=\text{CH}_2$</p> <p>a. 1,5-octadiene</p> <p>b. 3,7-octadiene</p> <p>c. 4,8-dioctene</p> | <p>7. $\text{CH}_3\text{CH}_2\text{C}\equiv\text{CCH}_2\text{CH}_3$</p> <p>a. 3-hexene</p> <p>b. 3-hexyne</p> <p>c. 3-pentyne</p> |
|---|--|

Use the IUPAC rules to name the following structures.

