

Matter—Properties and Changes

Section 3.1 Properties of Matter

In your textbook, read about physical properties and chemical properties of matter.

Use each of the terms below just once to complete the passage.

chemical	mass	physical
density	properties	substance

Matter is anything with **(1)** _____ and volume. A **(2)** _____ is a form of matter with a uniform and unchanging composition. Substances have specific, unchanging **(3)** _____ that can be observed. Substances have both physical and chemical properties. **(4)** _____ properties can be observed without changing a substance’s chemical composition. Color, hardness, and **(5)** _____ are examples. Other properties cannot be observed without changing the composition of a substance. These are called **(6)** _____ properties. An example is the tendency of iron to form rust when exposed to air.

Label each property as either *physical* or *chemical*.

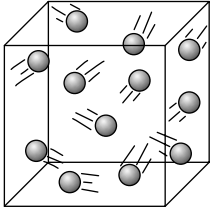
- _____ **7.** Chemical formula H₂O
- _____ **8.** Forms green carbonate when exposed to moist air
- _____ **9.** Remains unchanged when in the presence of nitrogen
- _____ **10.** Colorless
- _____ **11.** Solid at normal temperatures and pressures
- _____ **12.** Ability to combine with another substance
- _____ **13.** Melting point
- _____ **14.** Liquid at normal temperatures and pressures
- _____ **15.** Boiling point is 100°C
- _____ **16.** Conducts electricity
- _____ **17.** Density is $\frac{1\text{g}}{\text{cm}^3}$

Section 3.1 *continued*

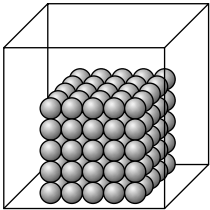
In your textbook, read about states of matter.

Label each drawing with one of these words: *solid, liquid, gas*.

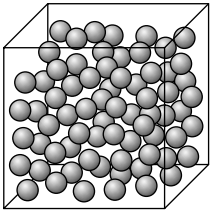
18.



19.



20.



For each statement below, write *true* or *false*.

- _____ 21. All matter that we encounter in everyday life exists in one of three physical forms.
- _____ 22. A solid has definite shape and volume.
- _____ 23. A liquid has a definite shape and takes on the volume of its container.
- _____ 24. A gas has both the shape and the volume of its container.
- _____ 25. The particles in a gas cannot be compressed into a smaller volume.
- _____ 26. Liquids tend to contract when heated.
- _____ 27. The particles in a solid are spaced far apart.
- _____ 28. The words *gas* and *vapor* can be used interchangeably.

Section 3.2 Changes in Matter

In your textbook, read about physical change and chemical change.

What kinds of changes do these words indicate? Write each word under the correct heading. Use each word only once.

boil	crumple	crush	explode
burn	ferment	freeze	grind
condense	melt	oxidize	rot
corrode	rust	tarnish	vaporize

Physical Change

Chemical Change

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____

9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____

For each item in Column A, write the letter of the matching item in Column B.

Column A

Column B

- _____ 17. The new substances that are formed in a chemical reaction
- _____ 18. A chemical reaction that involves one or more substances changing into new substances
- _____ 19. Shows the relationship between the reactants and products in a chemical reaction
- _____ 20. States that mass is neither created nor destroyed in any process
- _____ 21. The starting substances in a chemical reaction

- a. chemical change
- b. reactants
- c. products
- d. chemical equation
- e. law of conservation of mass

Answer the following question. Write an equation showing conservation of mass of reactants and products.

22. In a laboratory, 178.8 g of water is separated into hydrogen gas and oxygen gas. The hydrogen gas has a mass of 20.0 g. What is the mass of the oxygen gas produced?

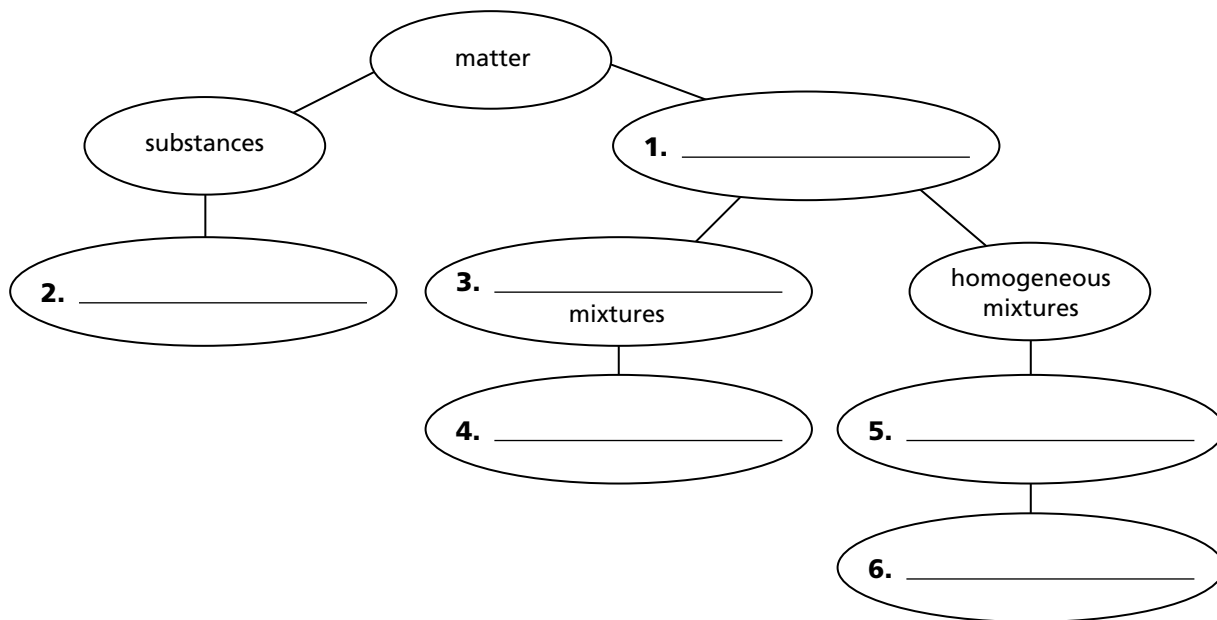
CHAPTER 3 STUDY GUIDE FOR CONTENT MASTERY

Section 3.3 Mixtures of Matter

In your textbook, read about pure substances and mixtures.

Use the words below to complete the concept map.

heterogeneous mixtures	salt–water mixture solutions	sand–water mixture water
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In your textbook, read about separating mixtures.

For each item in Column A, write the letter of the matching item in Column B.

Column A

- _____ 7. Separates substances on the basis of the boiling points of the substances
- _____ 8. Separates by formation of solid, pure particles from a solution
- _____ 9. Separates substances based on their movement through a special paper
- _____ 10. Separates solids from liquids by using a porous barrier

Column B

- a. filtration
- b. distillation
- c. crystallization
- d. chromatography

Section 3.4 Elements and Compounds

In your textbook, read about elements and compounds.

Circle the letter of the choice that best completes the statement or answers the question.

1. A substance that cannot be separated into simpler substances by physical or chemical means is a(n)
 - a. compound.
 - b. mixture.
 - c. element.
 - d. period.
2. A chemical combination of two or more different elements is a(n)
 - a. solution.
 - b. compound.
 - c. element.
 - d. period.
3. Which of the following is an example of an element?
 - a. water
 - b. air
 - c. sugar
 - d. oxygen
4. Which of the following is an example of a compound?
 - a. gold
 - b. silver
 - c. aspirin
 - d. copper
5. What are the horizontal rows in the periodic table called?
 - a. block elements
 - b. groups or families
 - c. grids
 - d. periods
6. What are the vertical columns in the periodic table called?
 - a. block elements
 - b. groups or families
 - c. grids
 - d. periods

Label each substance as either an *element* or a *compound*.

- | | |
|---|-----------------------------------|
| _____ 7. silicon
_____ 8. sodium chloride
_____ 9. francium | _____ 10. nickel
_____ 11. ice |
|---|-----------------------------------|

Write the symbol for each element. Use the periodic table on pages 72–73 in your textbook if you need help.

- | | |
|---|--|
| _____ 12. neon
_____ 13. calcium
_____ 14. iron | _____ 15. titanium
_____ 16. fluorine |
|---|--|

In your textbook, read about the law of definite proportions.

Use the law of definite proportions and the equation below to answer the questions.

The law of definite proportions states that regardless of the amount, a compound is always composed of the same elements in the same proportion by mass.

$$\text{Mass percentage of an element (\%)} = \frac{\text{mass of element}}{\text{mass of compound}} \times 100\%$$

17. A 20.0-g sample of sucrose contains 8.4 g of carbon. What is the mass percentage of carbon in sucrose? Show your work.

Section 3.4 *continued*

- 18.** Sucrose is 51.50% oxygen. How many grams of oxygen are in 20.0 g of sucrose? Show your work.

- 19.** A 2-g sample of sucrose is 6.50% hydrogen. What is the mass percentage of hydrogen in 300 g of sucrose? Explain your reasoning.

- 20.** Two compound samples are found to have the same mass percentages of the same elements. What can you conclude about the two samples?

In your textbook, read about the law of multiple proportions.

Use the law of multiple proportions to answer the questions and complete the table below.

The law of multiple proportions states that if the elements X and Y form two compounds, the different masses of Y that combine with a fixed mass of X can be expressed as a ratio of small whole numbers.

- 21.** Two compound samples are composed of the same elements, but in different proportions. What can you conclude about the two samples?

For each compound in the table, fill in the ratio of the mass of oxygen to the mass of hydrogen.

Compound	Mass of Oxygen	Mass of Hydrogen	Mass O/Mass H
H ₂ O	16 g	2 g	22.
H ₂ O ₂	32 g	2 g	23.

- 24.** Write a brief statement comparing the two mass ratios from the table.

- 25.** Are H₂O and H₂O₂ the same compound? Explain your answer.
